

## Corrigendum

### Removal of reactive azo dye using platinum-coated titanium electrodes with the electro-oxidation process\*

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The original version of the above article was published with errors in the order of Figures 1 to 6. The correct order of figures is as below.

The author apologizes for any confusion caused. The original article has been updated.

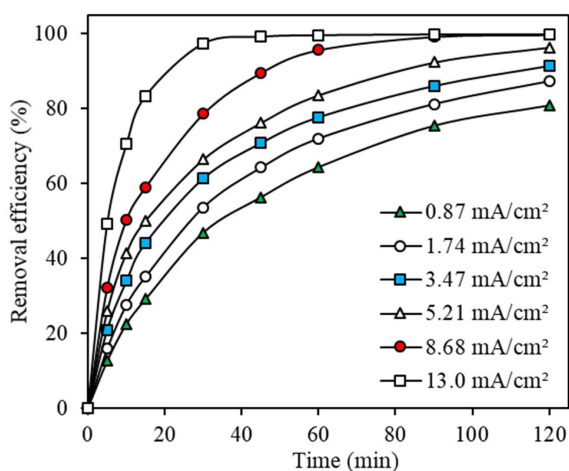


Fig. 1. Effect of current density on dye removal ( $C_0 = 100$  mg/L, pH= 7,  $T = 20^\circ\text{C}$ , Electrolyte= 0.5 g/L NaCl).

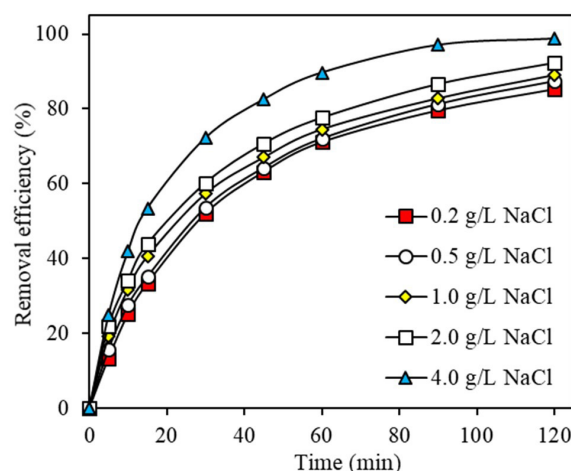


Fig. 2. Effect of NaCl concentration on dye removal ( $C_0 = 100$  mg/L, pH= 7,  $T = 20^\circ\text{C}$ ,  $J = 1.74$  mA/cm<sup>2</sup>).

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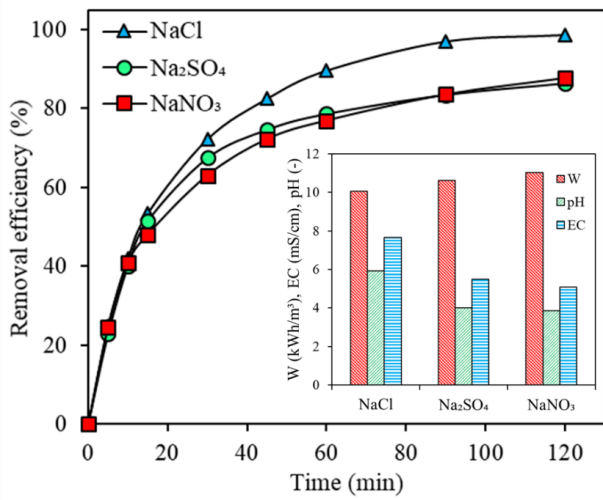


Fig. 3. Effect of electrolyte type on dye removal ( $C_0 = 100$  mg/L, pH = 7,  $T = 20^\circ\text{C}$ ,  $J = 1.74$  mA/cm<sup>2</sup>, Electrolyte = 4 g/L).

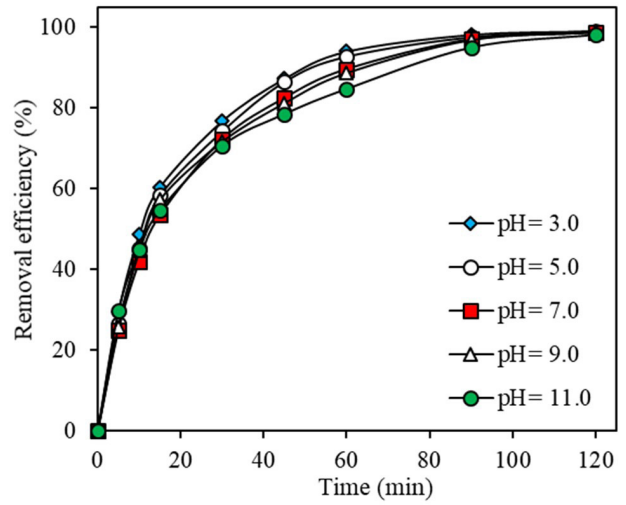


Fig. 4. Effect of pH on dye removal ( $C_0 = 100$  mg/L,  $T = 20^\circ\text{C}$ , Electrolyte = 4 g/L NaCl,  $J = 1.74$  mA/cm<sup>2</sup>).

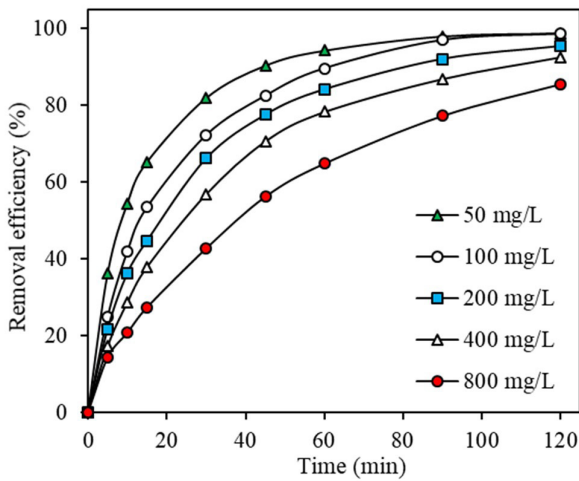


Fig. 5. Effect of initial dye concentration on dye removal (pH = 7,  $T = 20^\circ\text{C}$ , Electrolyte = 4 g/L NaCl,  $J = 1.74$  mA/cm<sup>2</sup>).

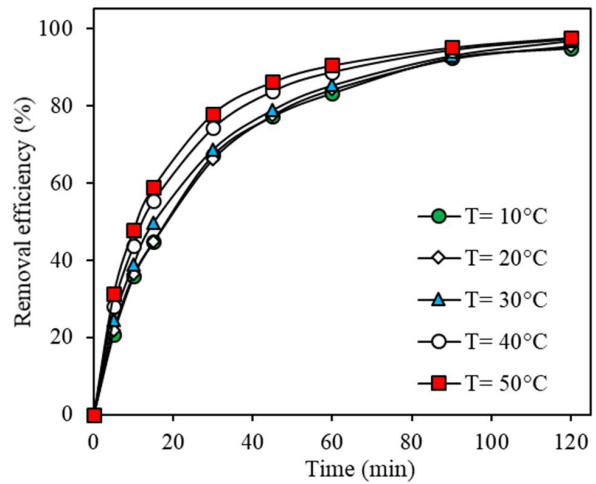


Fig. 6. Effect of temperature on dye removal (pH = 7,  $C_0 = 200$  mg/L, Electrolyte = 4 g/L NaCl,  $J = 1.74$  mA/cm<sup>2</sup>).